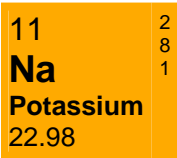
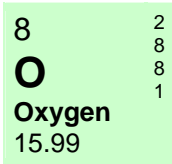


Topic 5 - The Periodic Table

- Mendeleev arranged the element cards into a 'solitaire-like' table. He played with them, by sorting and arranging the elements in many different combinations. He was able to identify gaps where elements, would be able to fit, that were ...
 - known to exist
 - not yet discovered
 - rare earth elements
 - identified by alchemists
 - In 1915 the Modern Periodic Table was reorganized, including more information about each element with a focus on ...
 - atomic structure
 - Chemical properties
 - Physical properties
 - reactivity rating
 - These are used worldwide and almost always use one or two letters that represent an element ...
 - Atomic Number
 - Mass Number
 - Atomic Symbol
 - Atomic Mass
 - Vertical columns form a **group** of elements (*numbered 1-18*) The horizontal rows (*numbered 1-7*) are called ...
 - lists
 - types
 - family
 - periods
5.  In this element – **Sodium** – 11 refers to the ...

 - mass
 - number
 - reactivity
 - ion charge

6.  In this element – **Oxygen** – the **numbers down the right side** indicate the ...

 - mass
 - toxicity
 - number
 - ion charge
- In the periodic table the following elements would be identified as the Noble Gases.
 - Be, Mg, Ca, Sr, Ba, Ra
 - Li, Na, K, Rb, Cs, Fr
 - He, Ne, Ar, Kr, Xe, Rn
 - Rf, Db, Sg, Bh, Hs, Mt, Uun
 - Sodium is identified in the excerpt from the periodic table above. How many neutrons does Sodium have?
 - 11
 - 12
 - 22
 - 23
 - As you move across the periodic table the properties of the elements change. The most reactive metals include ...
 - sodium and lithium
 - iron and copper
 - aluminum and carbon
 - lead and zinc